



# Citrate detection in Urine by Colorimetric Paper-based Sensor for Screening of Prostate cancer

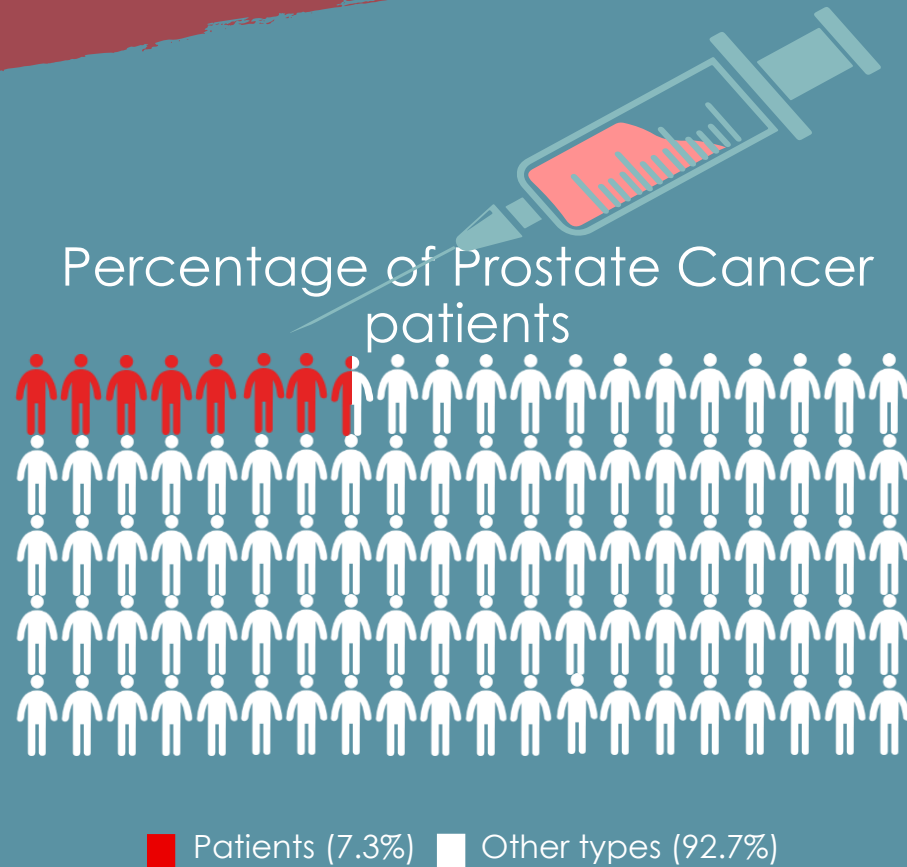


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# INTRODUCTION

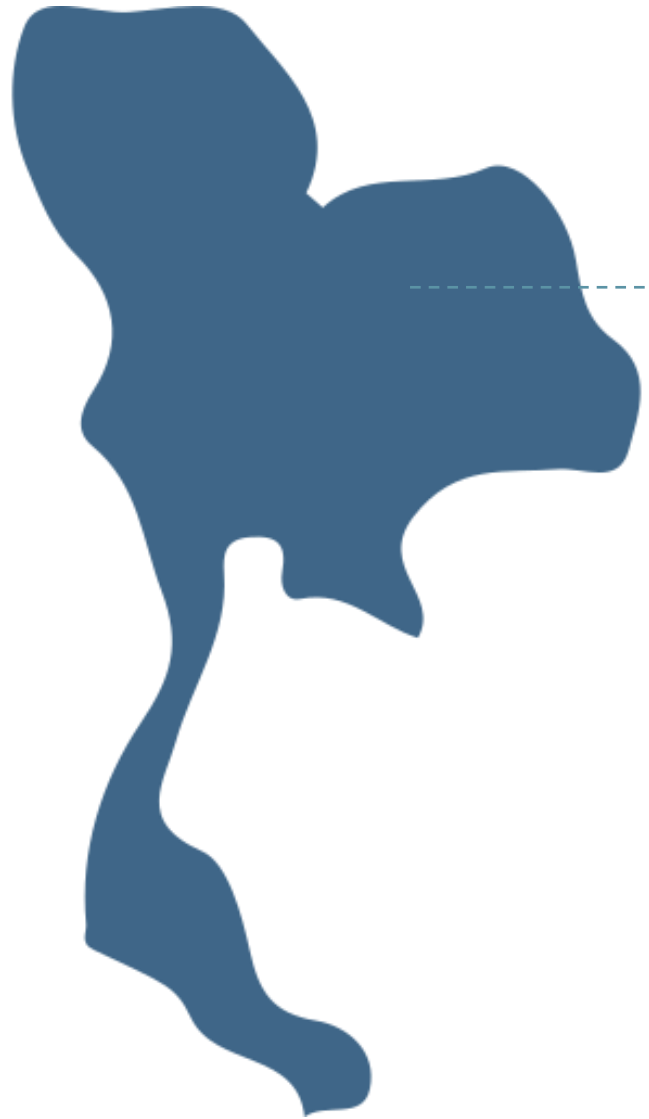


2018



(National Cancer Institute,  
Bangkok, Thailand, 2018)

# Introduction



4<sup>th</sup> most common cancer in Thailand



20-50% Survival rate (5 years cumulating)



50% of Men over 80 years old



gradual increase



# Symptoms of prostate cancer

✓ pee more frequently, often during the night



✓ rush to the bathroom



✓ difficulty in starting to pee (hesitancy)

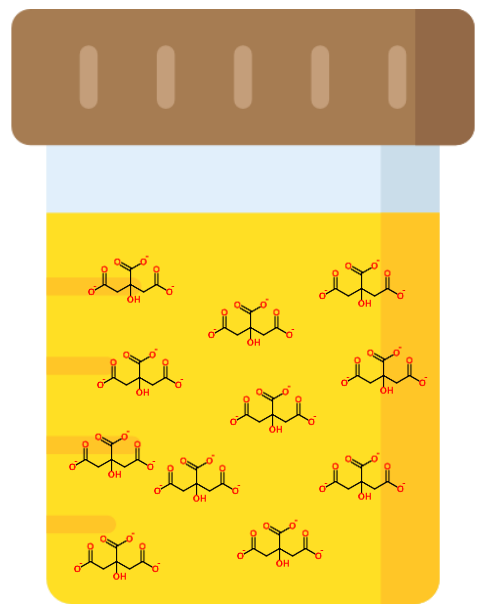


✓ strain or take a long time while peeing

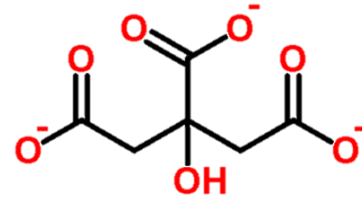




# Citrate in Urine



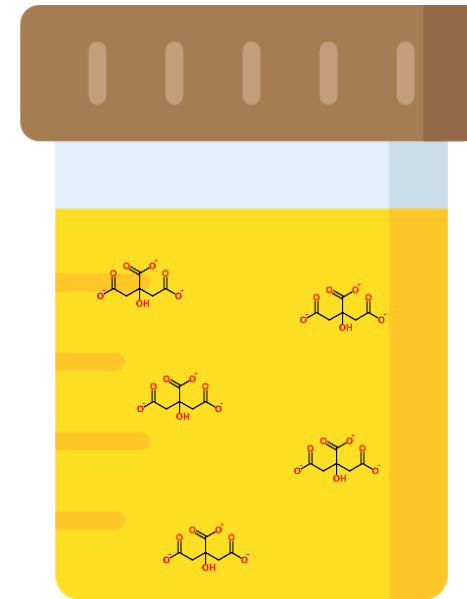
**9.5 mM**



Citrate



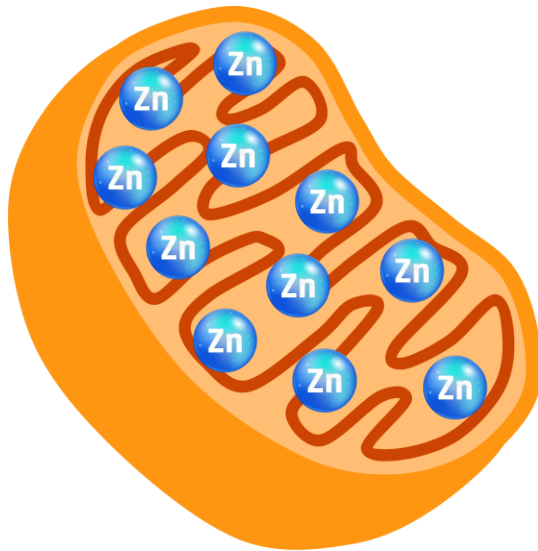
Prostate cancer



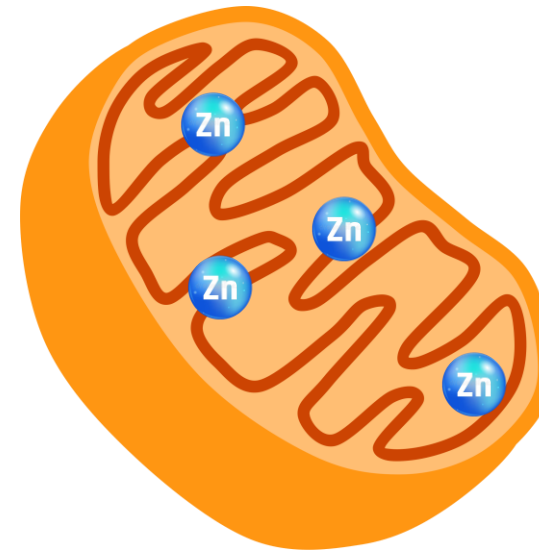
**6.9–2.4 mM**



# A decrease in the level of citrate

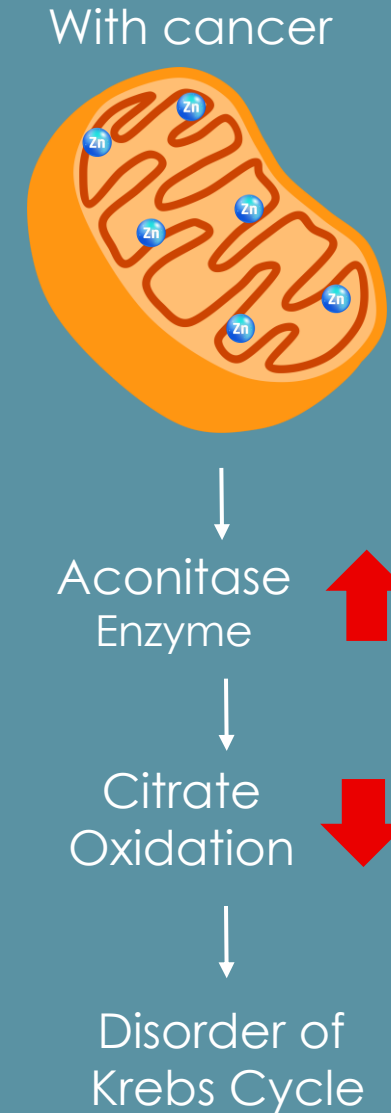
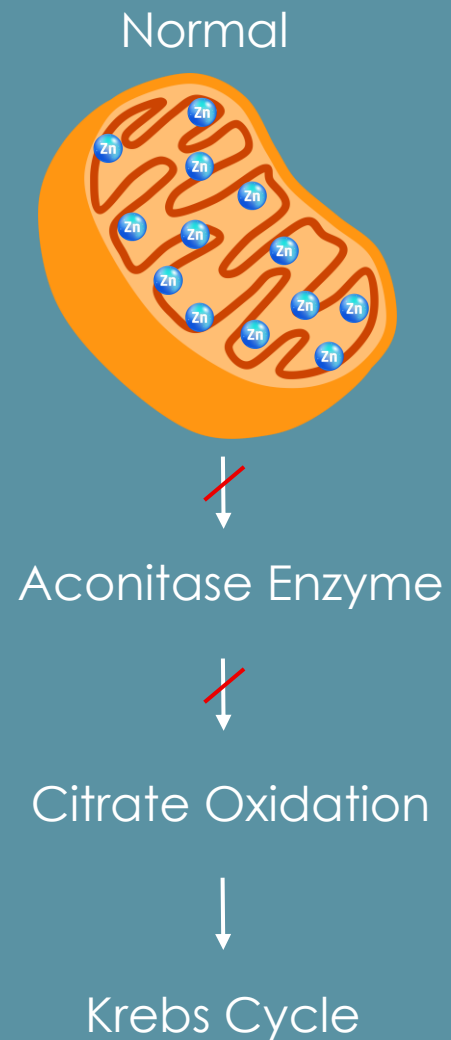


Normal

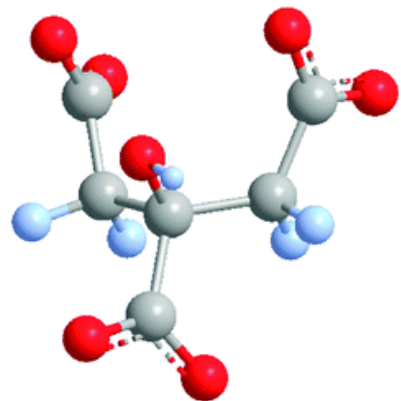
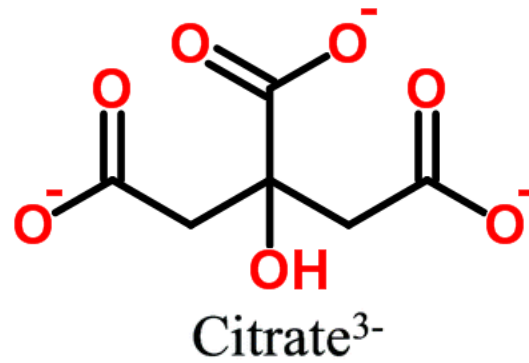


With cancer

# A decrease in the level of citrate

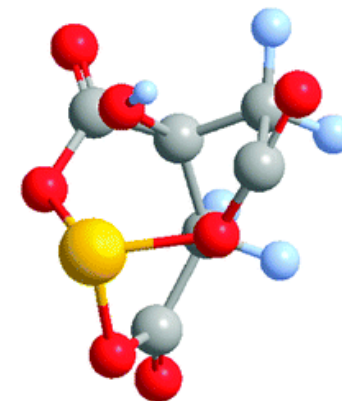
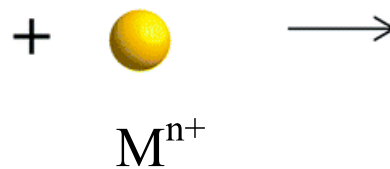


# Color from Complexes



(● C ● O ● H)

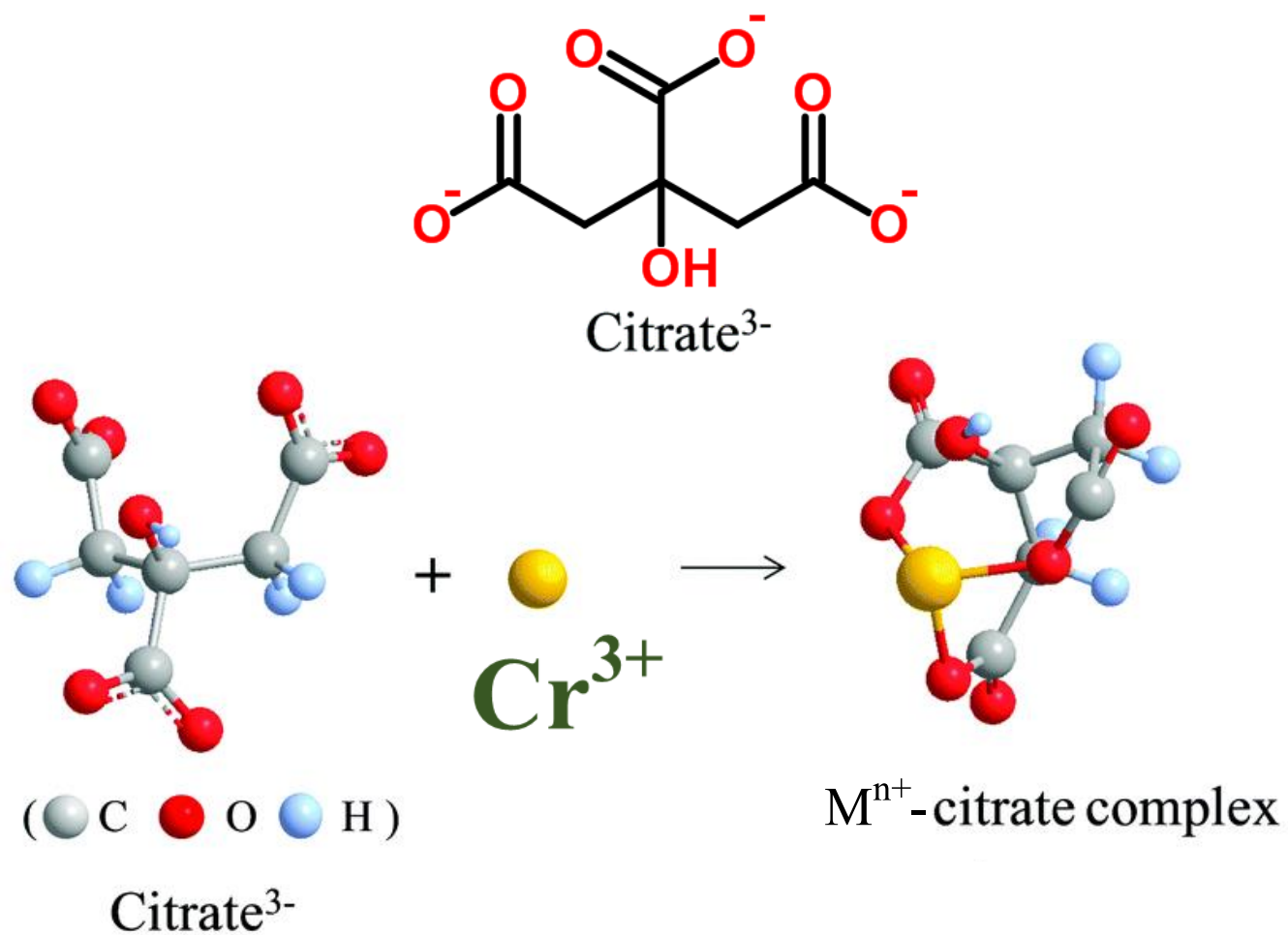
Citrate<sup>3-</sup>



$M^{n+}$  - citrate complex



# CITRATE COMPLEXS WITH METALS





# Finding Mole Ratio of the compound

$\text{Cr}(\text{NO}_3)_3$  10 mM



?  $\mu\text{L}$



Citrate 10 mM



?  $\mu\text{L}$

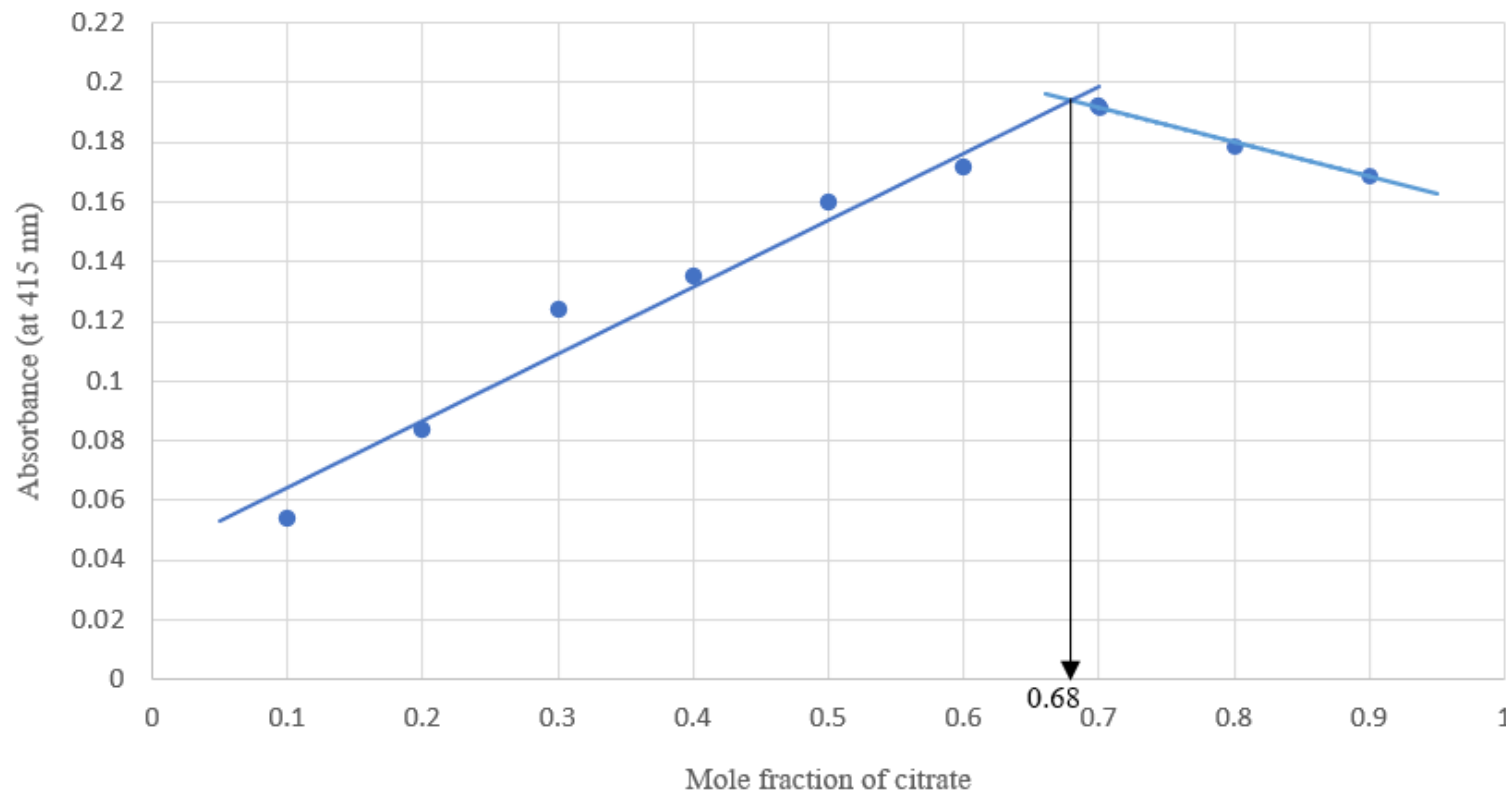
Solutions	Volume of Citrate 10 mM ( $\mu\text{L}$ )	Volume of $\text{Cr}^{3+}$ 10 mM ( $\mu\text{L}$ )
1	300	2700
2	600	2400
3	900	2100
4	1200	1800
5	1500	1500
6	1800	1200
7	2100	900
8	2400	600
9	2700	300



# Finding Mole Ratio of the compound

## Job's plot

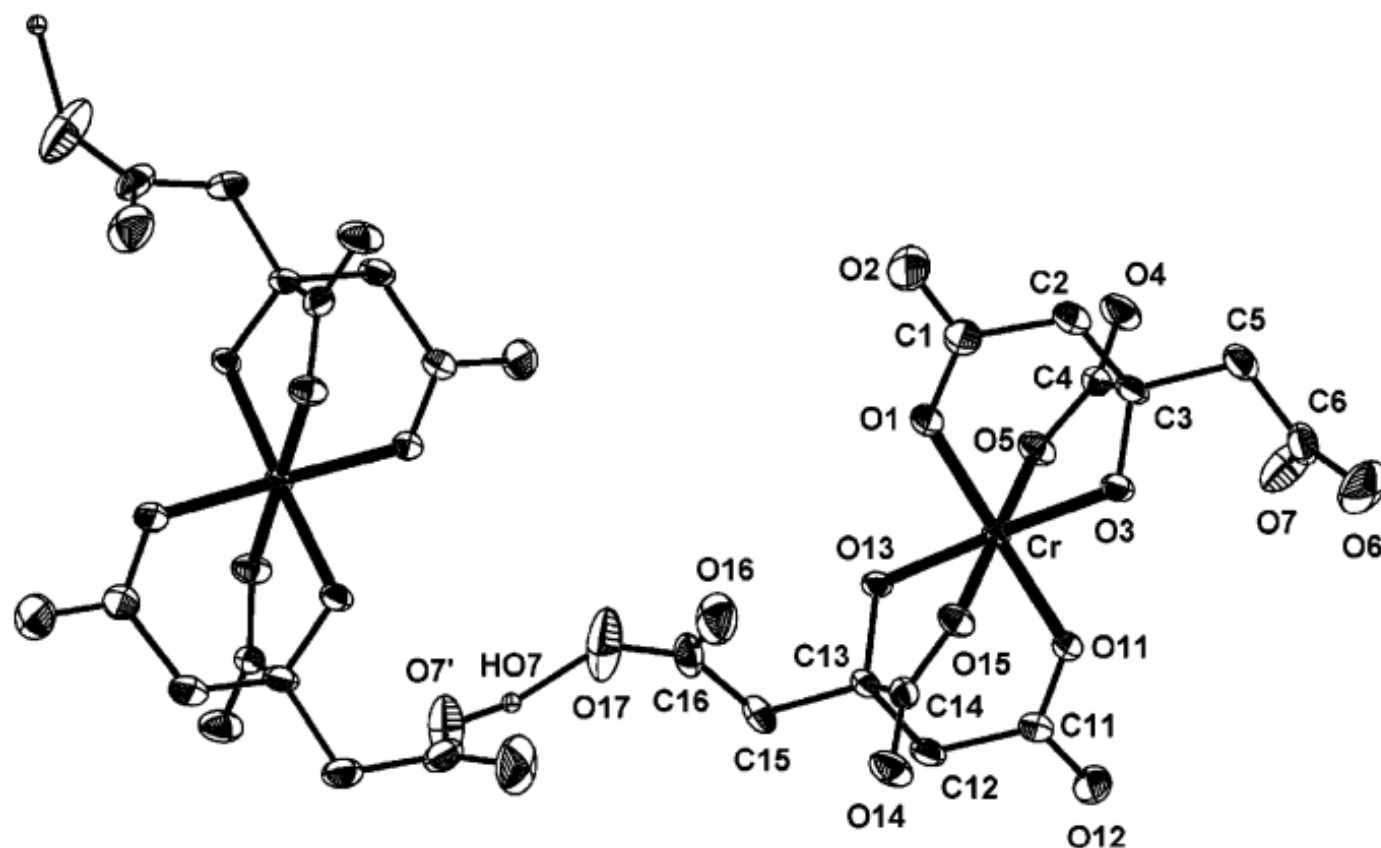
Job's plot for determining the stoichiometry of the complex between  $\text{Cr}^{3+}$  and citrate



$\text{Cr}^{3+}$  : Citrate  
1 : 2



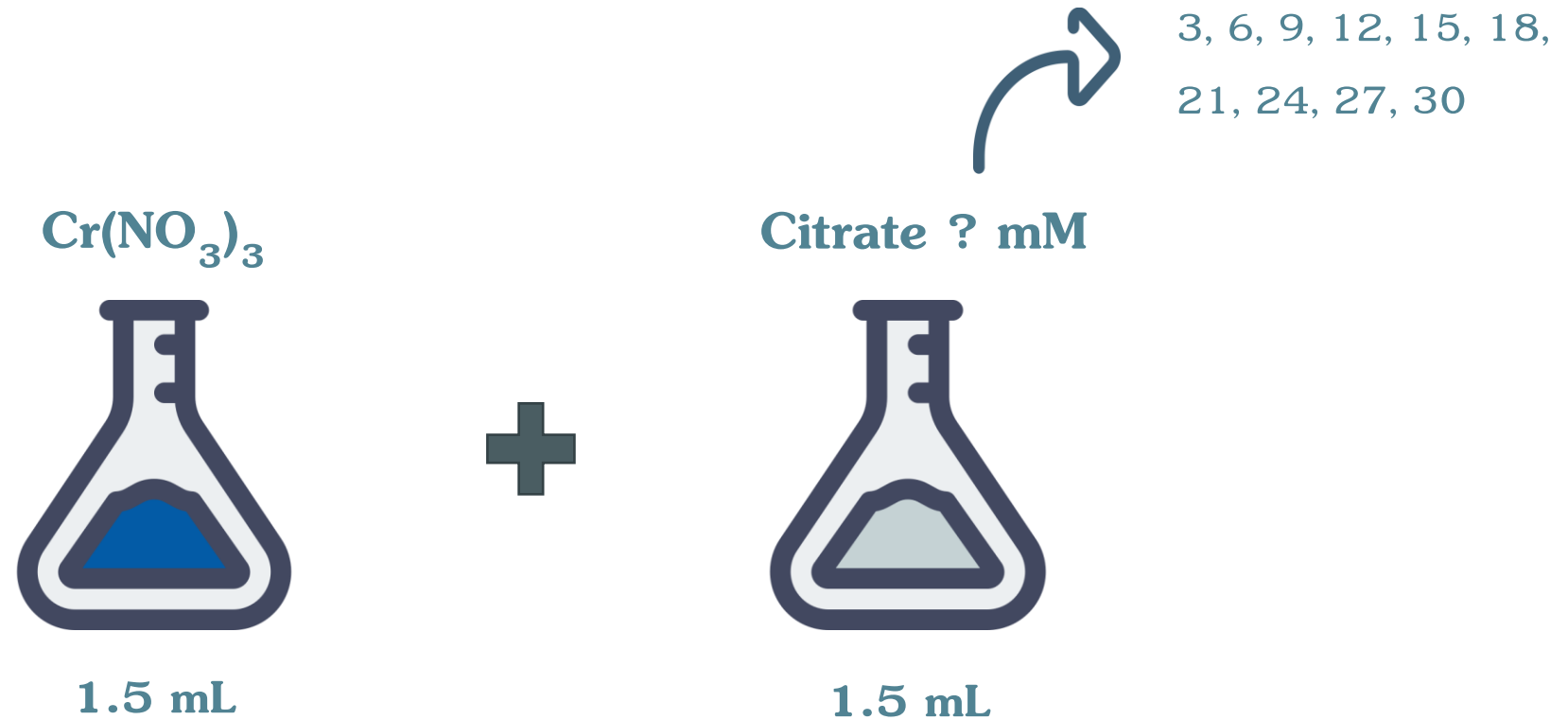
# Finding Mole Ratio of the compound



ORTEP structure of the  
[Cr(C<sub>6</sub>H<sub>4</sub>O<sub>7</sub>)(C<sub>6</sub>H<sub>5</sub>O<sub>7</sub>)] 4-anion



# Testing varied concentration of citrate in DI water

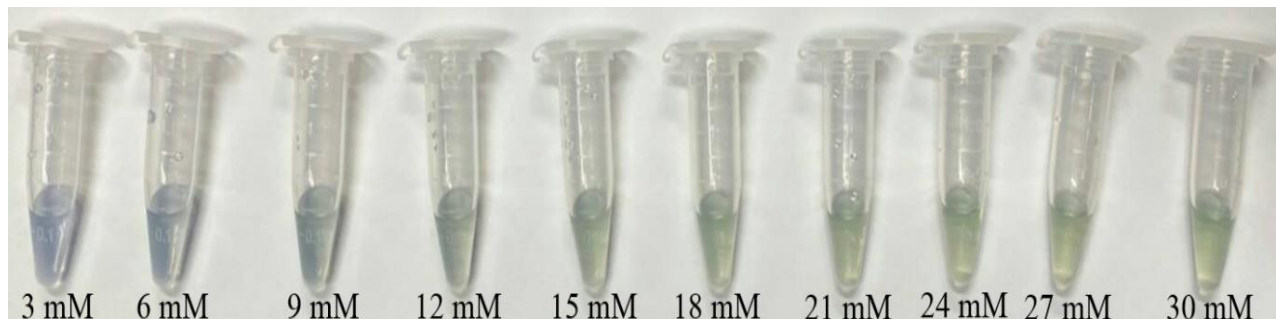
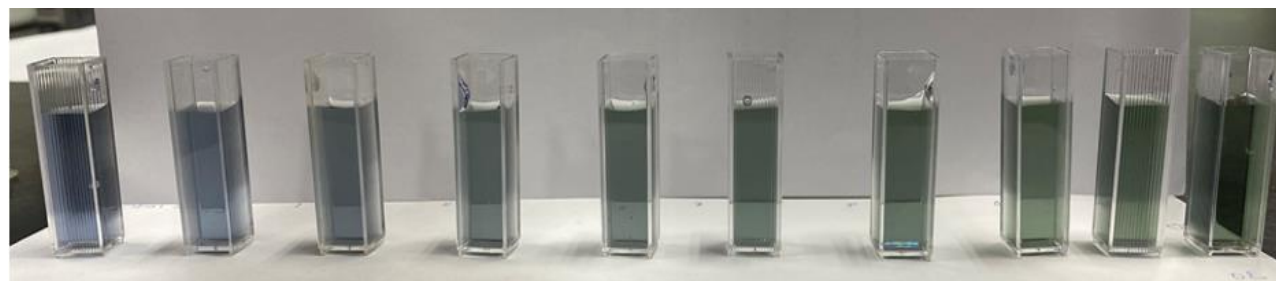




# Testing varied concentration of citrate in DI water

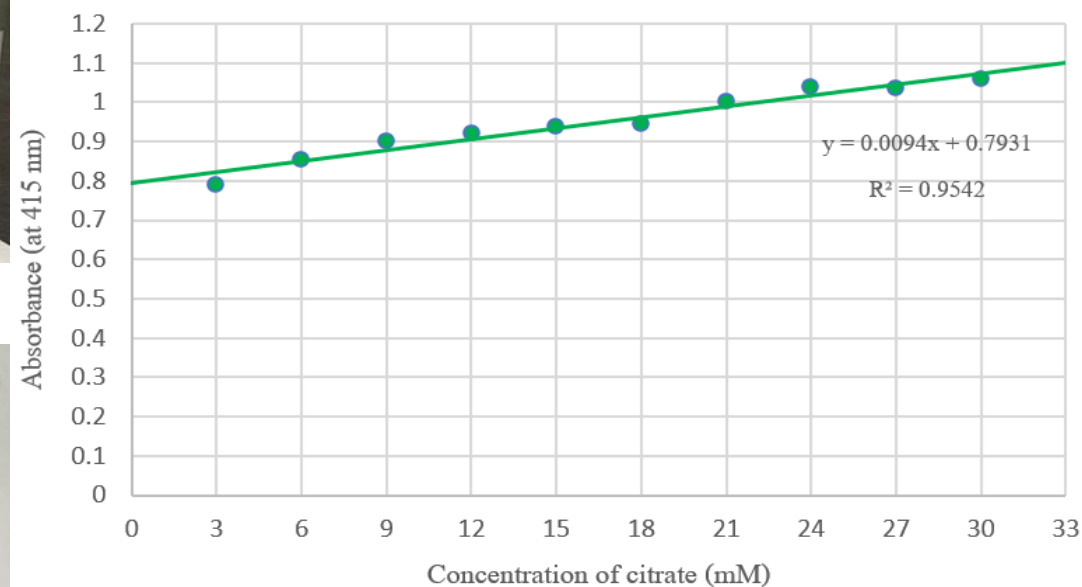


## Experiment in DI water



**Citrate Concentration**

Calibration curve of citrate in Cr(III)





# Preparation of Artificial Urine

(0.650 g)  $\text{Na}_3\text{C}_6\text{H}_5\text{O}_7$

(4.600 g)  $\text{NaCl}$

(0.651 g)  $\text{CaCl}_2$

(0.651 g)  $\text{MgCl}_2$

(1.600 g)  $\text{KCl}$

(2.800 g)  $\text{KH}_2\text{PO}_4$

$\text{C}_4\text{H}_9\text{N}_3\text{O}_2$  (1.100 g)

$\text{CO}(\text{NH}_2)_2$  (25.000 g)

$\text{Na}_2\text{SO}_4$  (2.300 g)

$\text{NH}_4\text{Cl}$  (1.000 g)

$\text{Na}_2\text{C}_2\text{O}_4$  (0.023 g)

**Artificial Urine 1 L**



# Experiments in Artificial Urine

Artificial Urine



Citrate ? mM

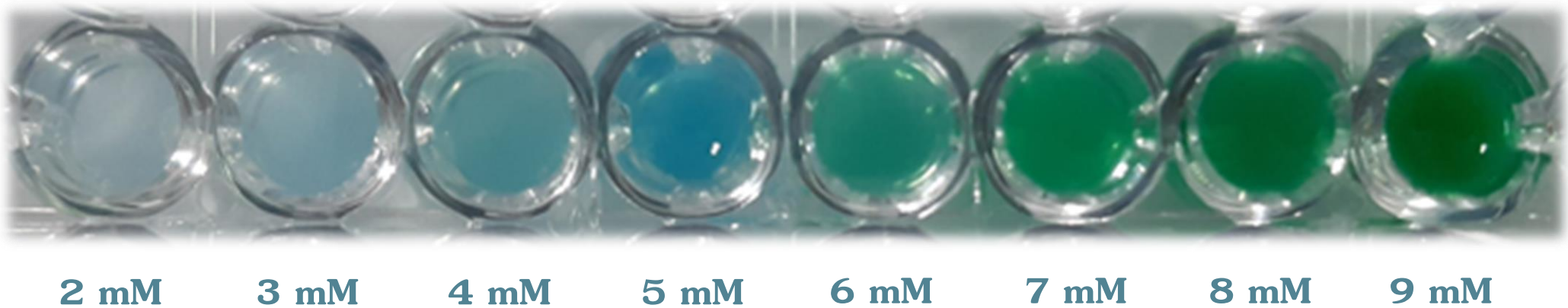


2, 3, 4, 5, 6, 7, 8, 9





# Experiments in Artificial Urine

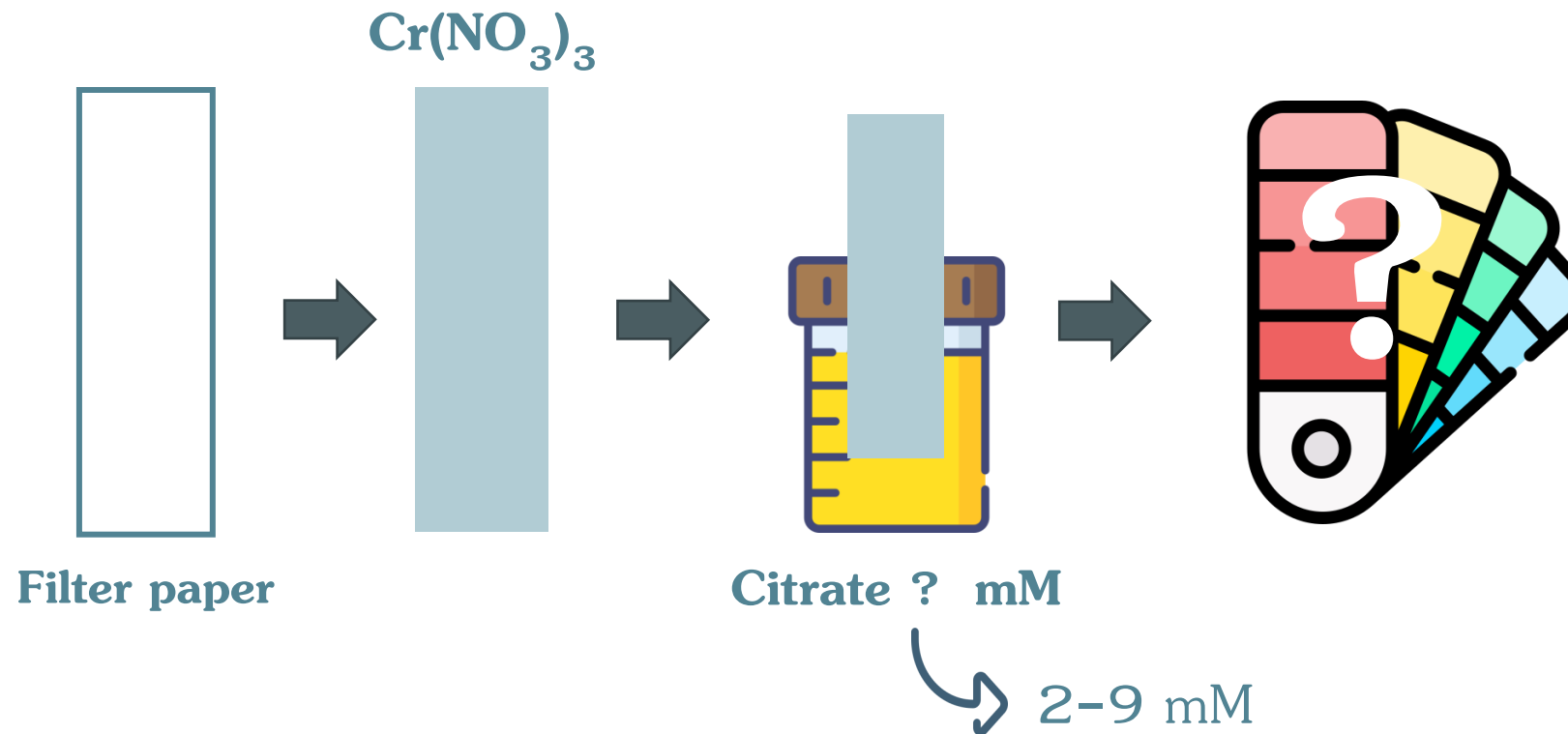


Citrate Concentration

Color change of chromium(III) ions and different concentration of citrate in artificial urine  
(The range covers the change of citrate concentration in human urine)

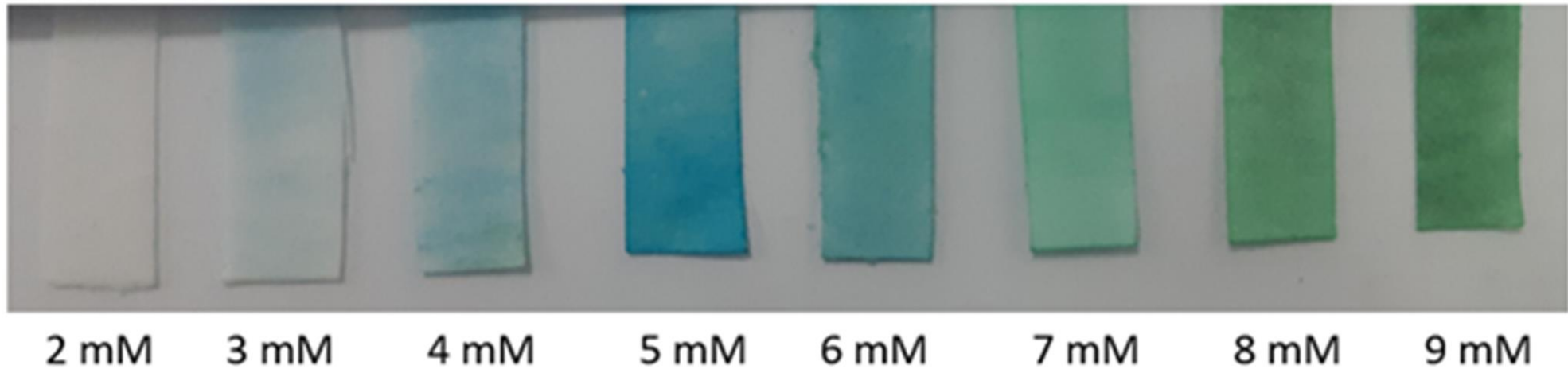


# Preparation of Paper-based Sensor





# Citrate Detection by Paper-based Sensor

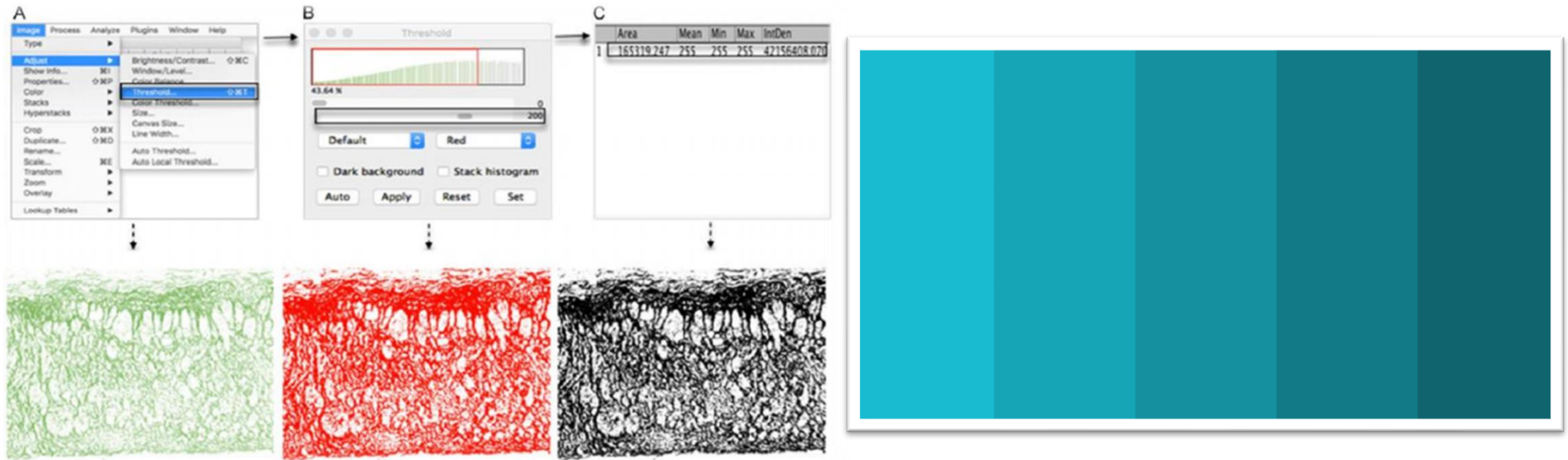


Low citrate



High citrate

# Further Development



- Using ImageJ program to find grey value and plot calibration curve
- Making a color chart for the colorimetric paper-based sensor for simple analysis and real usage



# References



- Costello, L. C., & Franklin, R. B. (1998). Novel Role of Zinc in the Regulation of Prostate Citrate Metabolism and Its Implications in Prostate Cancer, *296*(December 1997), 285–296.
- Gautier-luneau, I., Merle, C., Phanon, D., Lebrun, C., Biaso, F., Serratrice, G., & Pierre, J. (2005). New Trends in the Chemistry of Iron ( iii ) Citrate Complexes : Correlations between X-ray Structures and Solution Species Probed by Electrospray Mass Spectrometry and Kinetics of Iron Uptake from Citrate by Iron Chelators, *2207–2219*. <https://doi.org/10.1002/chem.200401087>
- Silva, R. O., Nabeshima, C. T., Bellini, M. H., & Courrol, C. (2012). Early Diagnosis of Prostate Cancer by Citrate Determination in Urine with Europium – Oxytetracycline Complex, *66*(8), 958–961. <https://doi.org/10.1366/11-06546>